

## Solubility of solids in solutions forming a complex ion

1) Solid sodium chloride is added to 5.00 L of  $0.100 \text{ mol L}^{-1}$  silver nitrate solution. Calculate the minimum mass of sodium chloride that would be needed to produce a saturated solution of AgCl. Assume that there is no change in volume when the sodium chloride is added.  $M(\text{NaCl}) = 58.5 \text{ g mol}^{-1}$

Discuss reasons for the fact that a precipitate of silver chloride dissolves on the addition of excess aqueous ammonia.